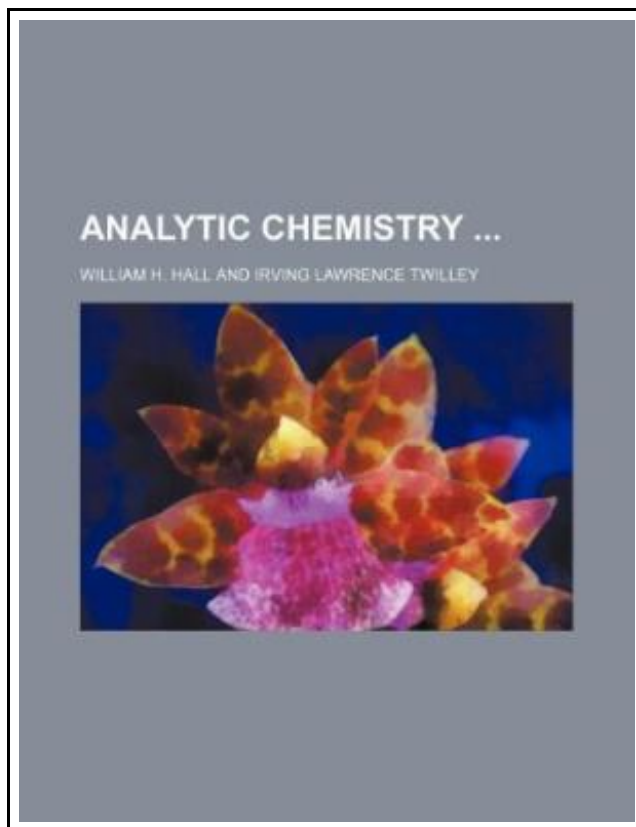


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Rarebookclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1920 Excerpt: .magnesium, and potassium. Give complete directions for their separation. 37. A solution contains aluminium chloride and zinc sulphate. How may the bases and acids be detected? 38. A sample powder consists of the sulphates of silver, magnesium and ferrous iron. Give complete directions for the separation of the bases, including equations wherever a precipitate is formed. 39. In what order are solvents applied to a powder? Name the solvents for silver chloride, cadmium sulphide, lead sulphate, lead chromate, aluminium hydroxide, and nickel sulphide. Describe the process in detail for dissolving barium sulphate. 40. Using formulas only, tabulate the commonly occurring bases into five groups, and the commonly occurring acids into three groups, by the use of general precipitants. 41. Give the complete tests for one member of each group of bases and for one member of each group of acids, mentioned in Question 40. TEST QUESTIONS S9 42. What is the function of carbon bisulphide in the tests for bromides and iodides? How may an iodide be removed from a solution that contains both an iodide and a bromide? What respective precautions should be taken in testing for chlorides, chlorates, cyanides, nitrates. phosphates and acetates? Give three tests for copper. 43. Make a complete set of equations for the interactions in the analysis of a salt composed of one base and one acid. With the aid of two equations, explain the changes in valence of an iron compound. Name the processes involved. 44. Using symbols, name four metals...



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